Cambridge International AS & A Level

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CHEMISTRY **9701/41**

Paper 4 A Level Structured Questions **October/November 2018**

MARK SCHEME

Maximum Mark: 100

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the October/November 2018 series for most Cambridge IGCSE™, Cambridge International A and AS Level components and some Cambridge O Level components.

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Generic Marking Principles

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

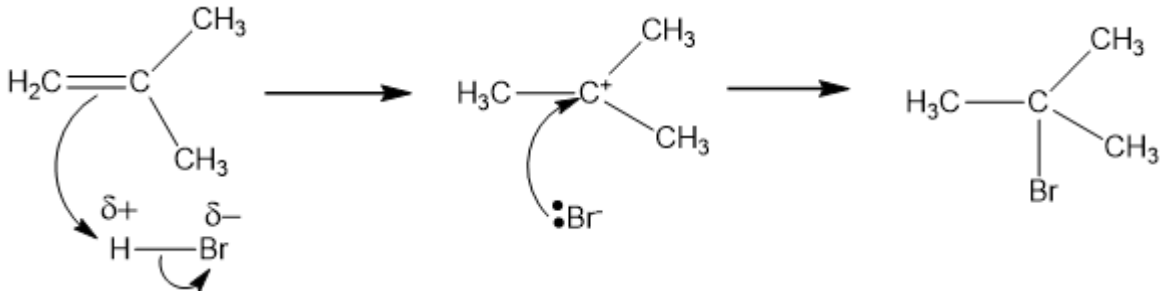
Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

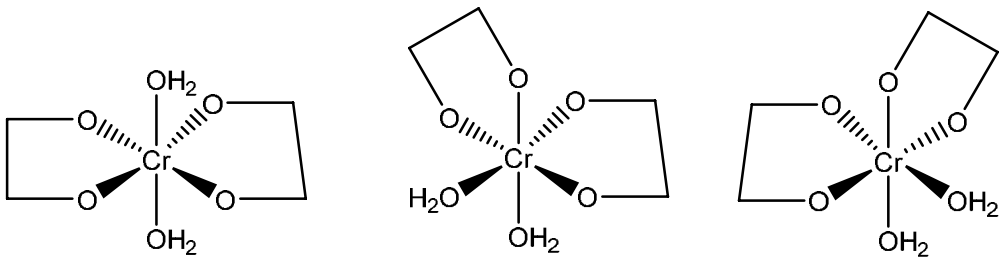
Rules must be applied consistently e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5: Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).
GENERIC MARKING PRINCIPLE 6: Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

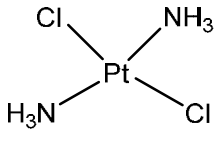
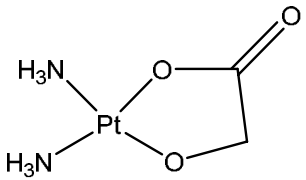
<table>
<tr><th>Question</th><th>Answer</th><th>Marks</th></tr>
<tr><td>1(a)(i)</td><td>
<table>
<tr><th>peak</th><th>organic compound</th><th>explanation</th></tr>
<tr><td>X</td><td>alkane</td><td>London forces only
OR no hydrogen bonding</td></tr>
<tr><td>Y</td><td>aldehyde</td><td>(Permanent dipole-dipole and London forces)</td></tr>
<tr><td>Z</td><td>carboxylic acid</td><td>(contains) hydrogen bonding</td></tr>
</table>
M1 peak assignments [1]
M2 explanation of Z OR X [1]
</td><td>2</td></tr>
<tr><td>1(a)(ii)</td><td>% of Z = 47/98 = 48%</td><td>1</td></tr>
<tr><td>1(b)(i)</td><td>^{37}Cl and ^{81}Br</td><td>1</td></tr>
<tr><td>1(b)(ii)</td><td>M peak $CH_2{}^{35}Cl^{79}Br$
M+2 peak $CH_2{}^{37}Cl^{79}Br$ OR $CH_2{}^{35}Cl^{81}Br$
M+4 peak $CH_2{}^{37}Cl^{81}Br$
two correct scores 1 mark
all 3 correct scores 2 marks</td><td>2</td></tr>
<tr><td>1(c)(i)</td><td> 

M1 correct dipole on HBr AND any correct curly arrow [1]
M2 two other correct curly arrows AND lone pair required on Br^- [1]
M3 intermediate [1]</td><td>3</td></tr>
</table>

Question	Answer	Marks
1(c)(ii)	(major product is) formed via the most **stable tertiary carbocation / intermediate** **OR tertiary** halogenoalkane formed via more **stable carbocation / intermediate**	**1**
1(d)(i)	**M1 ratio** of the concentrations of **solute** in two (immiscible) solvents [1] **M2** at equilibrium [1]	**2**
1(d)(ii)	$K_{partition}$ = (x/10)/(1.25-x/50) [1] 4.75(1.25-x) = 5x x = 5.9375/9.75 = **0.61 g** [1] correct answer scores [2]	**2**

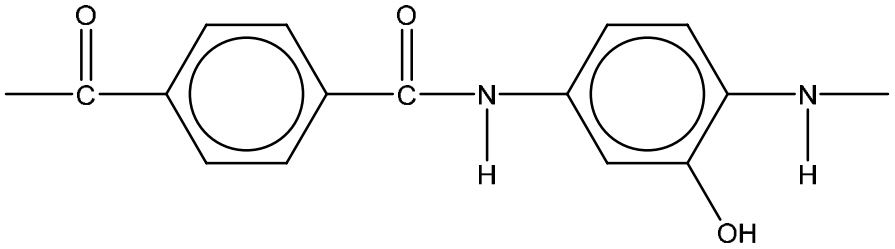
Question	Answer	Marks
2(a)	species that forms **dative bond**(s) to a (central) **metal atom / ion**	**1**
2(b)	 any two structures [1] × 2	**2**
2(c)(i)	$K_{sp} = [Ca^{2+}][C_2O_4^{2-}]$ [1] units $mol^2\ dm^{-6}$ [1]	**2**
2(c)(ii)	$[Ca^{2+}] = [C_2O_4^{2-}] = 6.65 \times 10^{-3}/128.1 = 5.19 \times 10^{-5}$ mol dm^{-3} [1] $K_{sp} = (5.19 \times 10^{-5})^2$ = **2.7×10^{-9}** $mol^2\ dm^{-6}$ [1]	**2**

Question	Answer	Marks
3(a)	[1] for each column number of unpaired electrons in element / 3d / 4s Cr / 5 / 1 Mn / 5 / 0 Fe / 4 / 0	**2**
3(b)	$2KMnO_4 \rightarrow K_2MnO_4 + O_2 + MnO_2$ formulae of K_2MnO_4 and O_2 [1] rest of the equation [1]	**2**
3(c)	**M1 d** orbitals split into two levels / lower and upper orbitals [1] **M2** visible light is **absorbed** and the complementary colour observed [1] **M3 electron**(s) promoted / excited [1]	**3**
3(d)(i)	precipitate **A** $[Cu(H_2O)_4(OH)_2]$ **OR** $Cu(OH)_2$ [1] solution **B** $[Cu(NH_3)_4(H_2O)_2]^{2+}$ [1] solution **C** $Cu(CH_3CO_2)_2$ [1]	**3**
3(d)(ii)	Na_2CO_3 **or** CO_3^{2-}	**1**
3(d)(iii)	$CuCO_3 + 2CH_3CO_2H \rightarrow Cu(CH_3CO_2)_2 + CO_2 + H_2O$	**1**
3(d)(iv)	any **two** for one mark • fizzing / bubbles / effervescence • solid **disappears** • green / blue solution (formed)	**1**

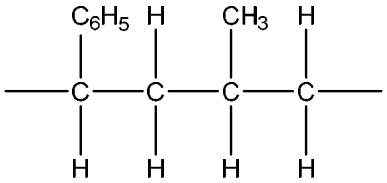
Question	Answer	Marks
3(e)	sum of the charges of the (four) ligands equals the oxidation number / charge of Pt **OR** a calculation Pt +2, NH_3 neutral / no charge, both Cl^-'s –1 (so no overall charge)	**1**
3(f)(i)	 [1] square planar **and** 180° [1]	**2**
3(f)(ii)	**M1** this can **bond / bind** with **DNA** [1] **M2** which prevents replication of the DNA / strand **OR** prevents cell division [1]	**2**
3(g)		**1**

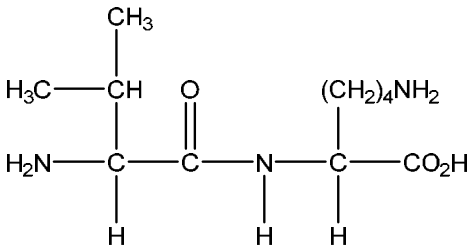
Question	Answer	Marks
4(a)	**M1** solubility decreases (down the Group) [1] **M2** because lattice energy and hydration energy **decreases** **OR** lattice energy and hydration energy become **less exothermic / more endothermic** [1] **M3** because hydration energy decreases to a **greater extent** (than does ΔH_{Latt}) [1]	**3**
4(b)(i)	(K_w =) $[H^+][OH^-]$	**1**

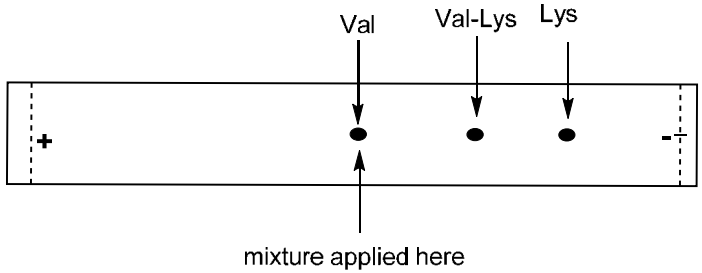
Question	Answer	Marks
4(b)(ii)	[1] or each correct tick effect of increasing temperature / decreases / stay the same / increase pH: ✓ (decreases) ratio of $[H^+]:[OH^-]$: ✓ (stay the same)	**2**
4(c)	$[H^+] = 10^{-13.25} = 5.62 \times 10^{-14}$ [1] $[OH^-] = K_w/[H^+] = 1.0 \times 10^{-14}/5.62 \times 10^{-14}$ $[OH^-]$ = **0.18** (0.178) (mol dm^{-3}) [1] ecf correct answer scores [2]	**2**
4(d)	$HCO_3^- + H^+ \rightarrow H_2CO_3$ **OR** $HCO_3^- + H^+ \rightarrow CO_2 + H_2O$ [1] $H_2CO_3 + OH^- \rightarrow HCO_3^- + H_2O$ [1]	**2**
4(e)(i)	***CH_3COOH*** + $H_2O \rightleftharpoons CH_3COO^- + H_3O^+$ [1] acid + base ⇌ base + acid [1]	**2**
4(e)(ii)	**M1** moles NaOH = $0.15 \times 20/1000$ = **0.0030** **AND** initial moles $CH_3COOH = 0.25 \times 30/1000$ **OR 0.0075** [1] **M2** equilibrium moles CH_3COOH = **0.0045 AND** equilibrium moles CH_3COONa = **0.0030** [1] **M3** $[CH_3COOH] = 0.0045/0.05$ = **0.090** AND $[CH_3COONa] = 0.003/0.05$ = **0.060** $[H^+] = K_a \times [CH_3COOH]/[CH_3COONa]$ = **2.625×10^{-5}** [1] **M4** pH = $-\log[H^+]$ = **4.6** [1] correct answer scores [4]	**4**
4(f)(i)	end point = 28 cm^3	**1**
4(f)(ii)	**M1** reaction **M** bromothymol (blue) / bromocresol (green) **AND** reaction **N** bromothymol (blue) / thymolphthalein [1] **M2** (both indicators have) a pH range / colour change **within** / **in** end-point / vertical region / sharp fall of the graph [1]	**2**

Question	Answer	Marks
5(a)(i)	[1] for each correct answer **F**: 3 peaks **G**: 6 peaks	**2**
5(a)(ii)	 one amide bond displayed in full [1] rest of the structure – one repeat unit only [1]	**2**
5(b)	[1] for each correct tick	**2**

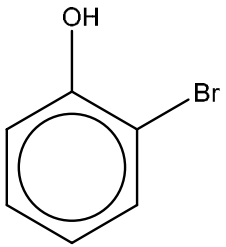
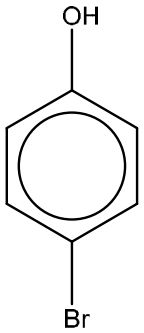
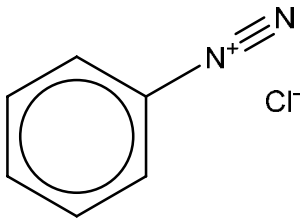
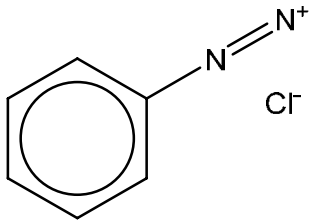
	σ-bonds only	π-bonds only	both σ- and π-bonds
bonds broken		✓	
bonds formed	✓		

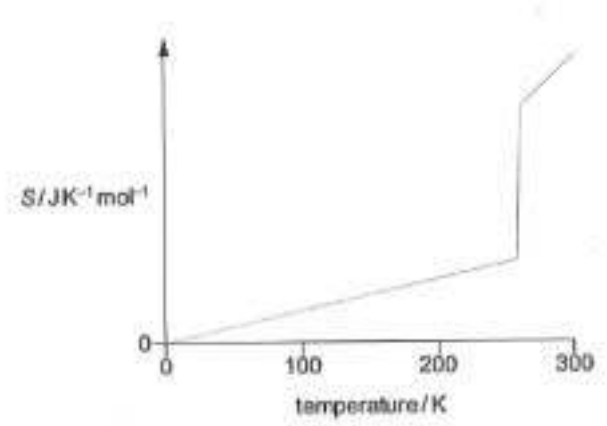
Question	Answer	Marks
5(c)	 **M1** length of chain with both monomers [1] **M2** continuation bonds [1]	**2**
5(d)(i)	C-C bonds are non-polar / have no dipole so cannot be hydrolysed [1]	**1**
5(d)(ii)	**M1** Hydrolysis using acid / base / alkali / enzymes [1] **M2** action of UV light [1]	**2**

Question	Answer	Marks
6(a)	 **M1** amide bond displayed [1] **M2** rest of the structure [1]	**2**

Question	Answer	Marks
6(b)	 **M1** valine on the cross [1] **M2** Val-Lys and Lys on the right of the cross (in any order) [1] **M3** relative order of Val-Lys and Lys (on the same side of the cross) [1] **Explanation** • Val does not move as it is a zwitterion / neutral (at pH6) **OR** Lys / Val-Lys move towards **negative** (pole) as they are **positively charged** • Lys moves the **furthest** as it has the **lowest** M_r (with the same positive charge) **OR** Val-Lys moves the **least** as it has the **largest** M_r (with the same positive charge) [1] × 2	**5**

Question	Answer	Marks
7(a)	**M1** C-X / C-C*l* / C-O bond is strong**er** (in chlorobenzene / phenol) [1] **M2** p-orbital / lone pair on C*l* / O(H) / X (in chlorobenzene / phenol) [1] **M3** electrons of the (C*l* / O / electronegative atom) **AND** overlap / delocalise with π-electron cloud / delocalise into ring [1]	**3**

Question	Answer	Marks
7(b)	 2-bromophenol  4-bromophenol structure and name correct [1]	**2**
7(c)(i)	step 1 **conc.** HNO_3 + H_2SO_4 (and temperare 50–55 °C) [1] step 2 Sn + HC*l* **AND one** of conc.HC*l* + heat [1] step 4 H_2O warm / heat [1]	**3**
7(c)(ii)	 **OR** 	**1**
7(c)(iii)	step 1 electrophilic substitution	**1**
7(c)(iv)	$C_6H_5NO_2$ + 6[H] → $C_6H_5NH_2$ + $2H_2O$	**1**

Question	Answer	Marks
8(a)	 **M1** continuous increase in *S* from 0–300 K (excluding m.p.) [1] **M2 steep** vertical increase in *S* **ONLY** at the m.p. **AND** continuous increase in *S* after m.p. [1]	**2**
8(b)	[1] for each correct tick <table><tr><td></td><td>negative $\Delta S^{\ominus}$</td><td>positive $\Delta S^{\ominus}$</td></tr><tr><td>solid dissolving in water</td><td></td><td>✓</td></tr><tr><td>water boiling to steam</td><td></td><td>✓</td></tr></table>	**1**
8(c)	$\Delta H^{\ominus}$ = (2 × C=O) + (3 × H-H) – (3 × C-H) – (C-O) – (3xO-H) $\Delta H^{\ominus}$ = (2 × 805) + (3 × 436) – (3 × 410)– (1 × 360) – (3 × 460) [1] $\Delta H^{\ominus}$ = 1610 + 1308 – 1230 – 360 – 1380 = **– 52** (kJ mol^{-1}) [1] ecf correct answer scores [2]	**2**
8(d)(i)	$\Delta S^{\ominus}$ = 127 + 70 – (214 + 3 × 131) [1] = – 410 (J K^{-1} mol^{-1}) [1] ecf correct answer scores [2]	**2**
8(d)(ii)	$\Delta G^{\ominus} = \Delta H^{\ominus} - T\Delta S^{\ominus}$ [1] $\Delta G^{\ominus}$ = –131 – (298 × –0.41) = **– 8.8(2)** (kJ mol^{-1}) [1] correct answer scores [2]	**2**

Question	Answer	Marks
8(d)(iii)	(as temperature increases) feasibility decreases	**1**
8(e)(i)	$2CH_3OH + 3O_2 \rightleftharpoons 2CO_2 + 4H_2O$ **OR** $2CH_3OH + 3O_2 \rightleftharpoons 2CO_2 + 4H^+ + 4OH^-$	**1**
8(e)(ii)	$E^{\ominus}_{cell}$ = 1.23 – 0.02 = **1.21** V	**1**